Where is the electron pair?

A QMC detective story

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Talk overview

What we will cover for sure and maybe cover...

Electrophilic amination: A QMC study

- The chemistry
- Electron-pair localization functions (original and modified)
- Sketch current and future work

Recently published: Amador et al., J. Chem. Phys. 126, 204308 (2007)

Electronic structure by quantum computation

- Polynomial algorithm for the exact solution of the Schrödinger equation on a quantum computer
- Recent results





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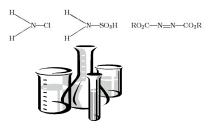
Making proteins: Synthesis of peptide bonds

Nucleophilic vs. electrophilic substitution

- \bullet Carbon-nitrogen bonds are usually formed by the nucleophilic attack of a nitrogen to an electrophilic carbon (S_N2)
- Reverse process: a nucleophilic carbon replaces a leaving group previously attached to an electrophilic nitrogen

$$X \longrightarrow C + HNRR' \longrightarrow C \longrightarrow NRR' \longleftarrow RR'N-X + C$$

Some reagents have been found and tested





Exploration using QMC

The mission

- Is it possible to rationalize the electronic properties of the already known reagents so as to be able to direct the search for even more effective ones?
- Approach the answer by:
 - Use DMC to calculate energetics
 - Calculate the electron-pair localization function (EPLF)
 rationalize values, find trends, explain differences, make predictions

Representative molecules

- Typical nucleophilic behavior:
 NH₃, NH₂-CH₃ (MA), NH-(CH₃)₂ (DMA), N-(CH₃)₃ (TMA)
- Possible electrophilic behavior: NH₂F, NH₂CF₃





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DMC calculations

Calculation parameters

- geometries optimized with GAMESS at MP2//631G* level
- cc-pVTZ basis set
- fourth order SMBH correlation function, absolute deviation minimization
- collection of 10⁶ walkers for EPLF
- ≈ 1 week computational time on 8 cores.
- All calculations carried out with the Zori code (http://www.zori-code.com)







Atomization enthalpy differences

Single-determinant energies

| Atomization enthalpy differences $ \Delta E_a^{ m calc} - \Delta E_a^{ m expt} $ (kJ/mol) | | | | | |
|---|--------|----------|-----|--|--|
| Method | | Molecule | | | |
| | NH_3 | MA | DMA | | |
| HF | 394 | 693 | 984 | | |

| | NH_3 | MA | DMA | TMA |
|----------------------|--------|------|---------|------|
| | 11113 | | D.W., C | |
| HF | 394 | 693 | 984 | 1278 |
| BLYP | 24 | 1 | 170 | 34 |
| B3LYP | 21 | 8 | 4 | 2 |
| B3PW91 | 6 | 1 | 2 | 3 |
| MP2FC | 29 | 58 | 70 | 75 |
| MP2FC//HF/6-31G* | 32 | 1121 | 76 | 86 |
| MP2FC//B3LYP/6-31G* | 32 | 1124 | 75 | 84 |
| MP2FC//MP2FC/6-31G* | 32 | 61 | 76 | 84 |
| DMC-SMBH//MP2/6-31G* | 3 | 9 | 14 | 91 |

Numbers are differences with the experimental value at 298.15 K.

Total energies are corrected by zero point energy (ZPE).



Electron Pair Localization Function

$$d_{\sigma\sigma}(\overrightarrow{r}) = \sum_{i=1}^{N} \langle \langle \delta(\overrightarrow{r} - \overrightarrow{r}_i) \min_{j:\sigma_j = \sigma_i} | \overrightarrow{r} - \overrightarrow{r}_j | \rangle \rangle,$$

$$d_{\sigma\bar{\sigma}}(\overrightarrow{r}) = \sum_{i=1}^{N} \langle \langle \delta(\overrightarrow{r} - \overrightarrow{r}_{i}) \min_{j:\sigma_{j} \neq \sigma_{i}} | \overrightarrow{r} - \overrightarrow{r}_{j} | \rangle \rangle.$$

EPLF

In terms of these average distances, the EPLF is defined as

$$EPLF(\overrightarrow{r}) = \frac{d_{\sigma\sigma}(\overrightarrow{r}) - d_{\sigma\bar{\sigma}}(\overrightarrow{r})}{d_{\sigma\sigma}(\overrightarrow{r}) + d_{\sigma\bar{\sigma}}(\overrightarrow{r})}.$$

EPLF range

(-1) unpaired (0) long-distance (1) Paired

A. Scemama, P. Chaquin, M. Caffarel. J. Chem. Phys., 121:1725-1735, 2004



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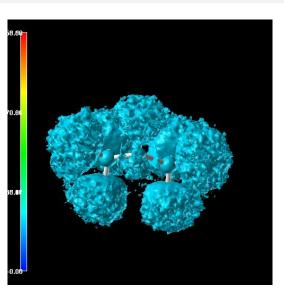
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Dimethylamine

A nucleophilic amine



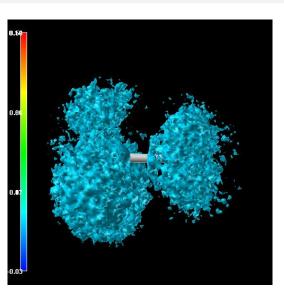
iso-EPLF surface for dimethylamine





Fluoroamine

An electrophilic amine

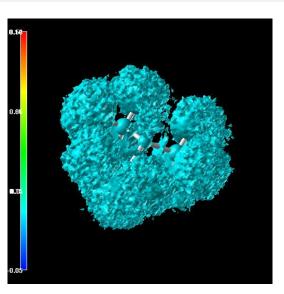


iso-EPLF surface for fluoroamine



Trimethylamine

A nucleophilic amine

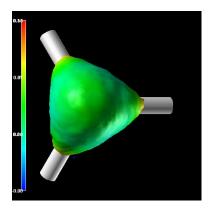


iso-EPLF surface for tri-methylamine





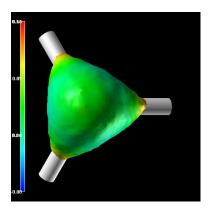
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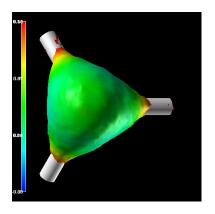
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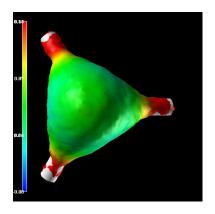
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EPLF projected on the electron density



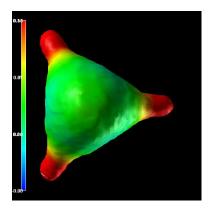
iso-density surface with projected color coded EPLF for ammonia



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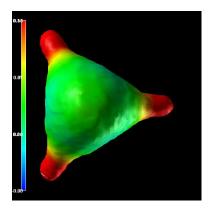
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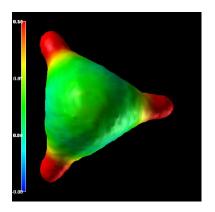
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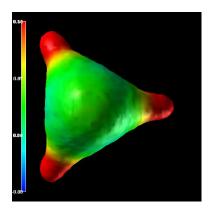
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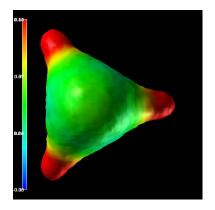
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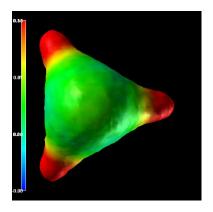
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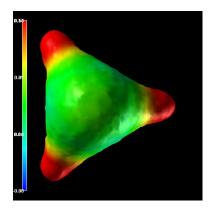
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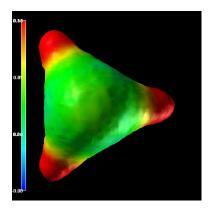
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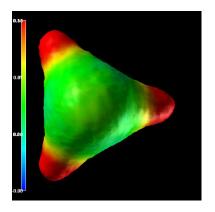
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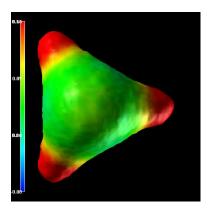
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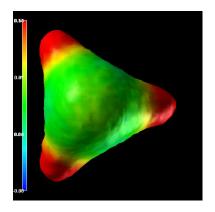
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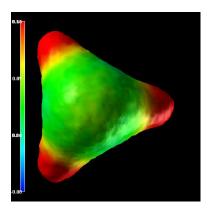
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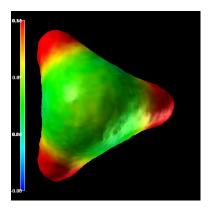
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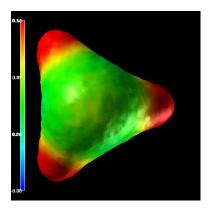
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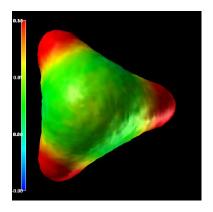
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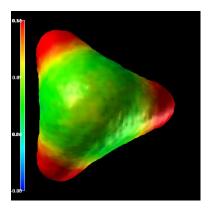
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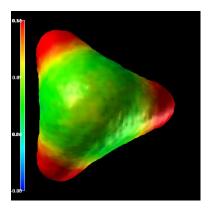
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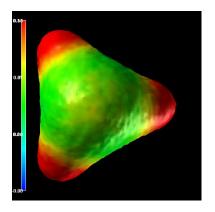
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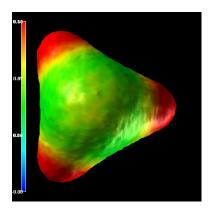
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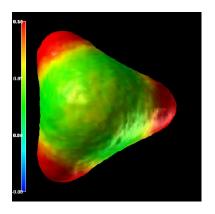
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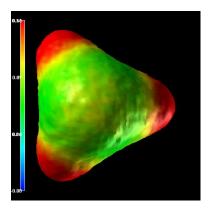
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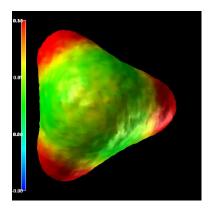
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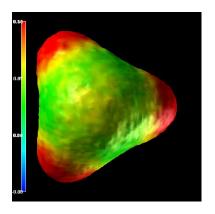
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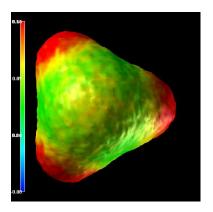
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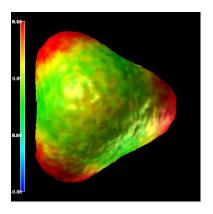
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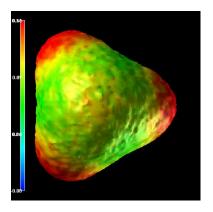
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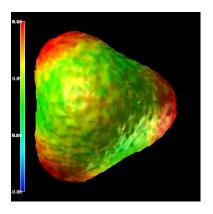
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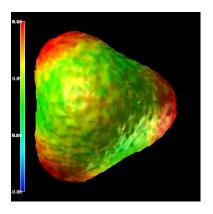
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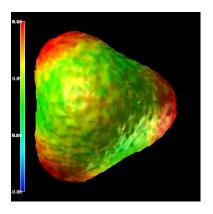
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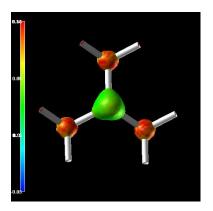
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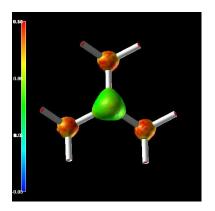
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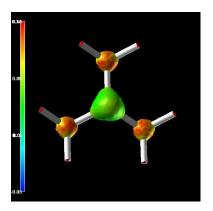
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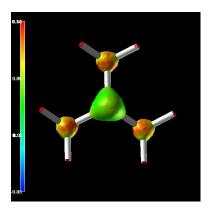
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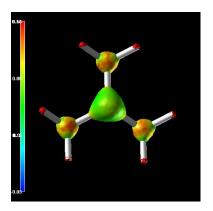
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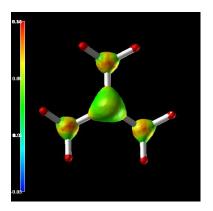
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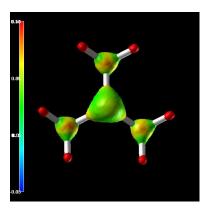
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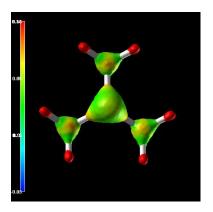
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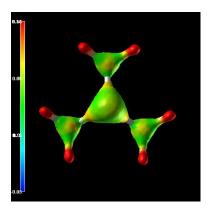
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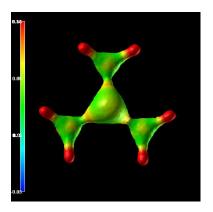
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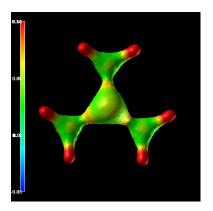
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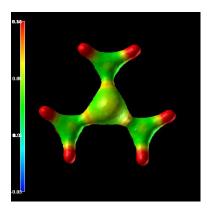
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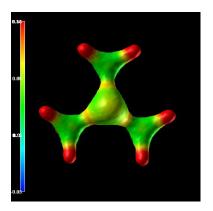
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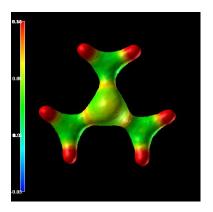
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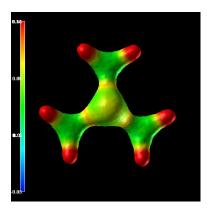
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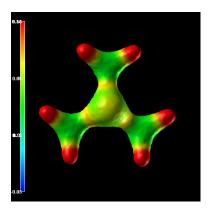
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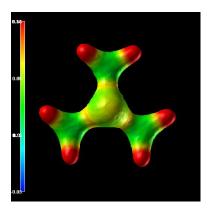
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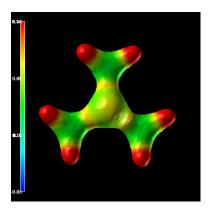
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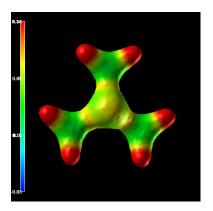
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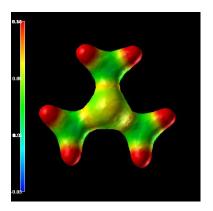
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EPLF projected on the electron density



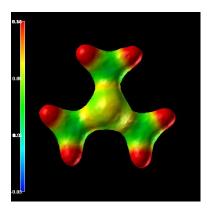
iso-density surface with projected color coded EPLF for trimethyl amine



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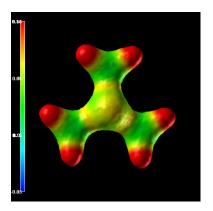
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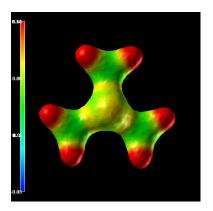
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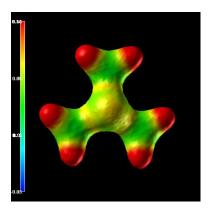
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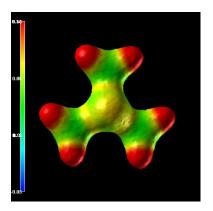
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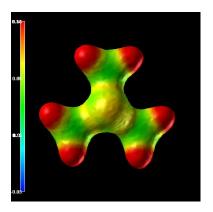
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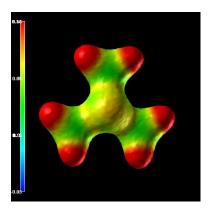
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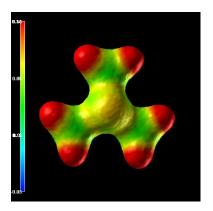
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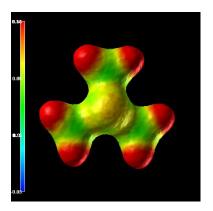
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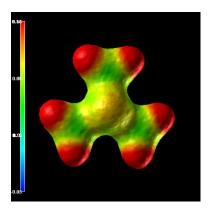
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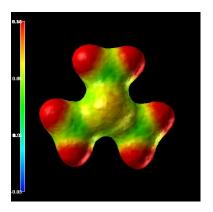
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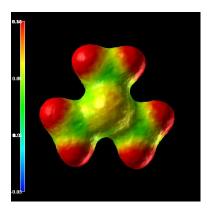
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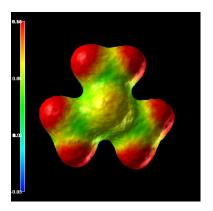
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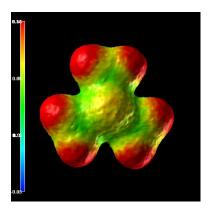
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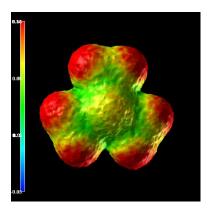
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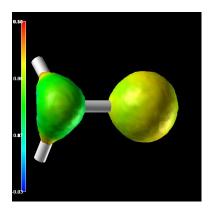
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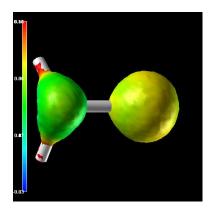
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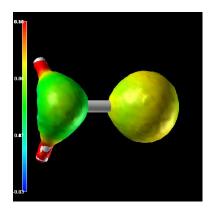
EPLF projected on the electron density







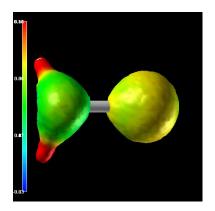
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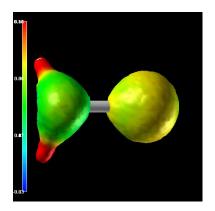
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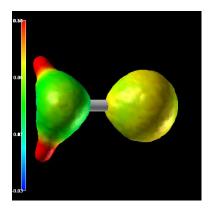
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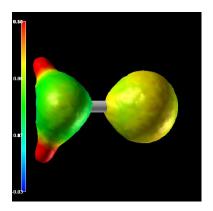
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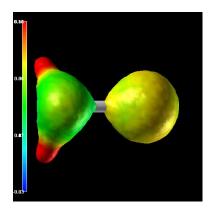
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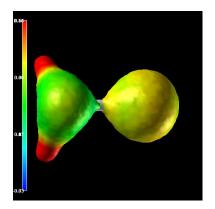
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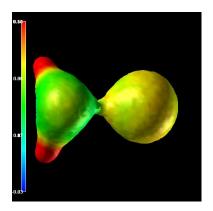
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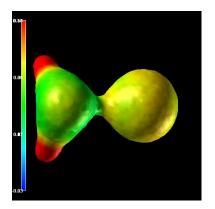
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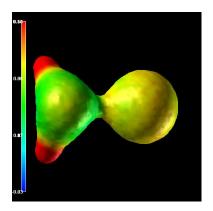
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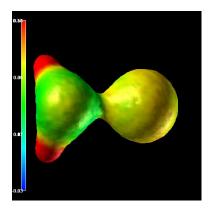
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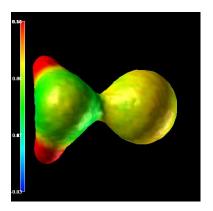
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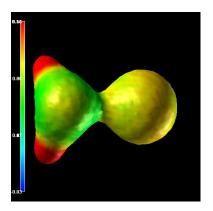
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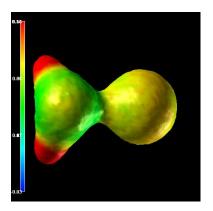
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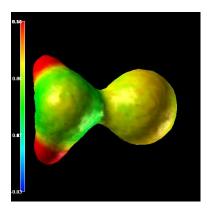
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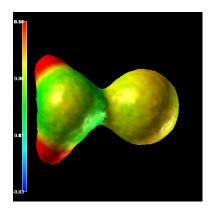
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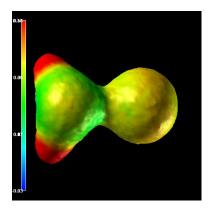
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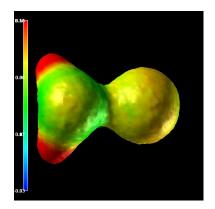
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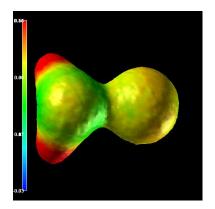
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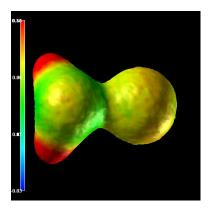
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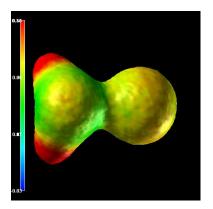
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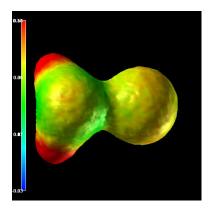
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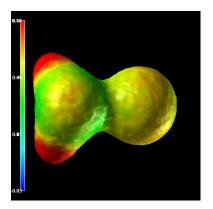
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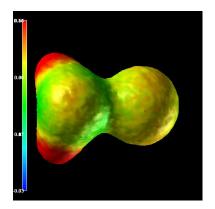
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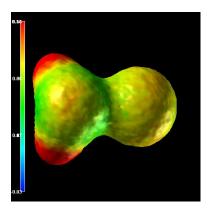
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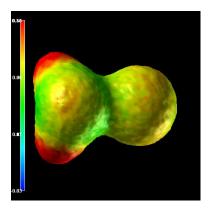
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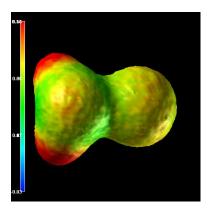
EPLF projected on the electron density







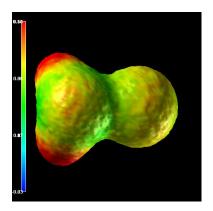
EPLF projected on the electron density







EPLF projected on the electron density



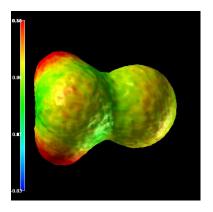
iso-density surface with projected color coded EPLF for fluoroamine



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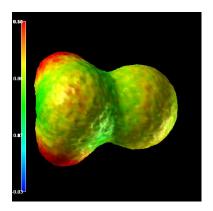
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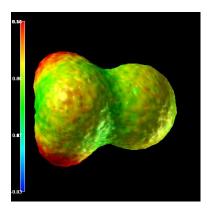
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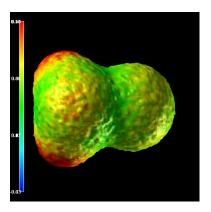
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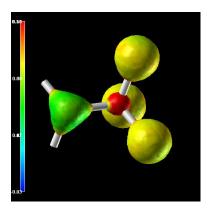
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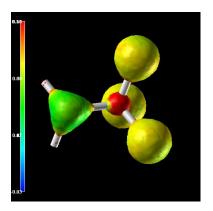
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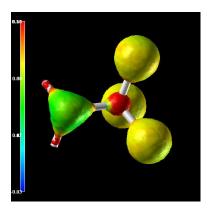
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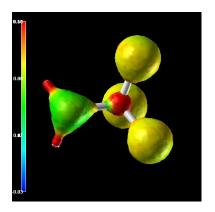
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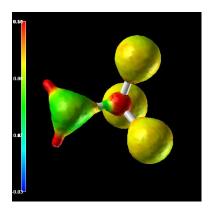
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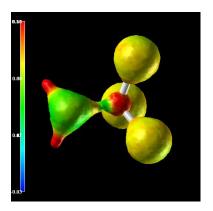
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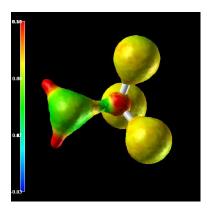
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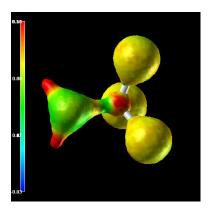
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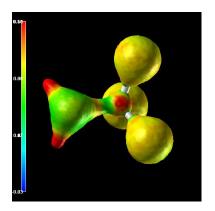
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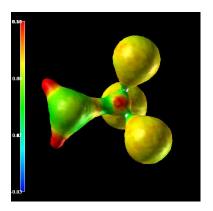
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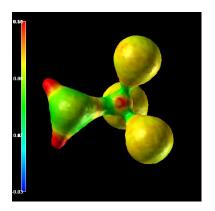
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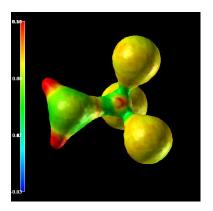
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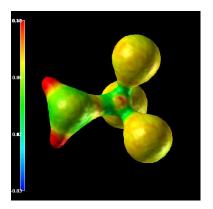
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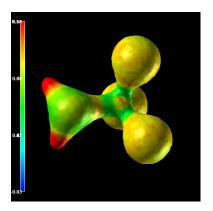
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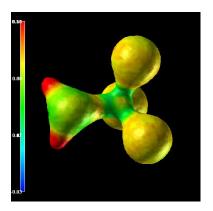
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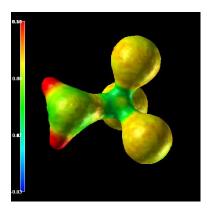
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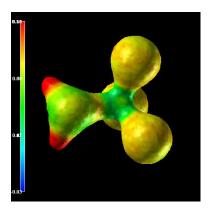
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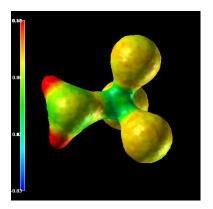
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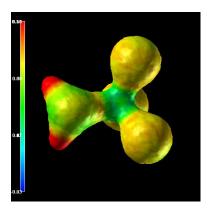
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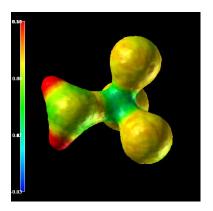
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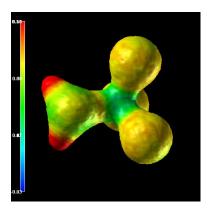
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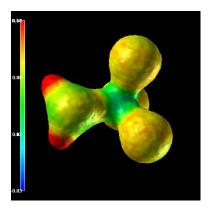
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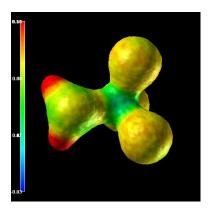
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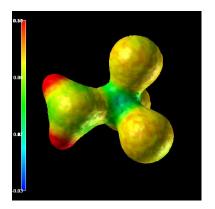
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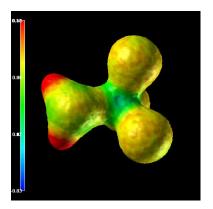
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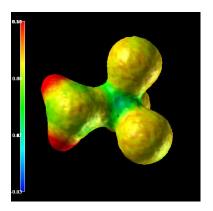
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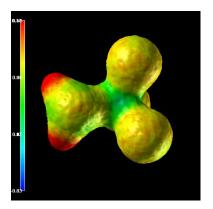
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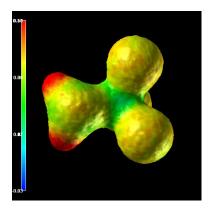
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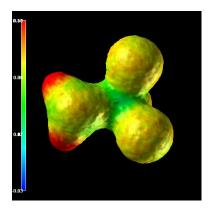
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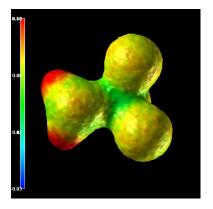
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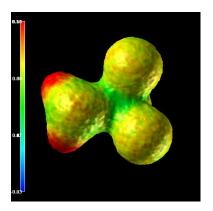
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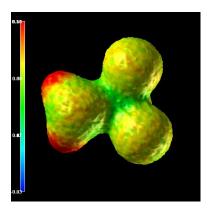
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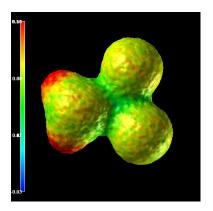
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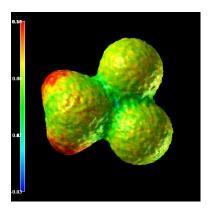
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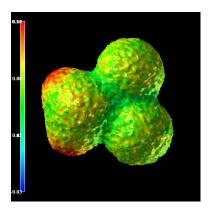
EPLF projected on the electron density







EPLF projected on the electron density







A conclusion

EPLF seems to be a very good hydrogen locator!

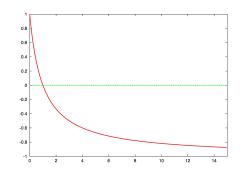
Let's go back to definition..

$$EPLF(\overrightarrow{r}) = \frac{d_{\sigma\sigma}(\overrightarrow{r}) - d_{\sigma\bar{\sigma}}(\overrightarrow{r})}{d_{\sigma\sigma}(\overrightarrow{r}) + d_{\sigma\bar{\sigma}}(\overrightarrow{r})}.$$

If the $\sigma\sigma$ and $\sigma\bar{\sigma}$ distances are proportional to each other, EPLF is insensitive to the inter-electronic distance

Assume
$$d_{\sigma\bar{\sigma}}(\overrightarrow{r}) = xd_{\sigma\sigma}(\overrightarrow{r})$$

$$EPLF(\overrightarrow{r}) = \frac{1-x}{1+x}$$





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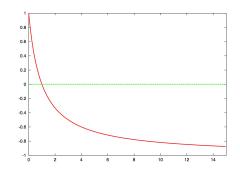
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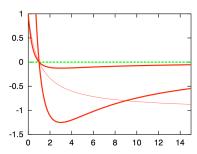


A distance-dependent variant of EPLF

For characterizing an **electro-nucleophilicity** scale, we are interested in measuring how *paired* and *localized* are electrons.

We introduce the EPLFN family of functions. In particular EPLF3, the EPLF density

$$EPLFN(\overrightarrow{r}) = \frac{d_{\sigma\sigma}(\overrightarrow{r}) - d_{\sigma\bar{\sigma}}(\overrightarrow{r})}{\left(d_{\sigma\sigma}(\overrightarrow{r}) + d_{\sigma\bar{\sigma}}(\overrightarrow{r})\right)^{N+1}}$$



But... where is the electron pair region?



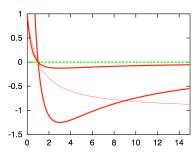


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But... where is the electron pair region?





Electron Localization Function (ELF)

The DFT original formula and inspiration for EPLF and EPLFN

Electron Localization Function (ELF)

Conditional pair probability for same spin —spherical average—,

$$P_{\text{cond}}^{\sigma\sigma}(\mathbf{r},s) = \frac{1}{3} \left[\sum_{j} |\nabla \phi_{j}|^{2} - \frac{1}{4} \frac{|\nabla \rho_{\sigma}|^{2}}{\rho_{\sigma}} \right] s^{2} + \dots$$

(smaller probability, more localization)
Define

$$D_{\sigma}(r) = \sum_{j} |\nabla \phi_{j}|^{2} - \frac{1}{4} \frac{|\nabla \rho_{\sigma}|^{2}}{\rho_{\sigma}};$$

and the corresponding quantity for the uniform electron gas

$$D_{\sigma}^{0}(r) = \frac{3}{5}(6\pi^{2})^{2/3}\rho_{\sigma}^{5/3}.$$





Electron Localization Function

Definition

Thus, ELF is defined as

$$ELF(r) = \frac{1}{1 + \left(D_{\sigma}/D_{\sigma}^{0}\right)^{2}}.$$

Interpretation

- ELF equals one when D_{σ} is zero —highest localization.
- ELF equals one half when the localization corresponds to that of the uniform electron gas.
- ELF is smaller for even smaller localization.

- A. D. Becke and K. E. Edgecombe. J. Chem. Phys. 92:5397-5403, 1990.
- B. Silvi and A. Savin. Nature 371:683-686, 1994.



EPLF in the ELF basin

Effect of electron correlation on the electron pair region density

For our purposes, the most important property of ELF is that defines a well-defined basin around the lone pair, *i. e.* defines the *pair's volume*.

| Descriptors of the electron pair basin | | | | | | | | |
|--|----------|-------|----------|--------|--|--|--|--|
| Molecule | Integral | | Distance | Volume | | | | |
| | LDA | QMC | | | | | | |
| TMA | 2.295 | 2.382 | 0.94 | 354 | | | | |
| DMA | 2.251 | 2.324 | 0.95 | 775 | | | | |
| MA | 2.215 | 2.232 | 0.95 | 1488 | | | | |
| NH_3 | 2.167 | 2.211 | 0.95 | 2294 | | | | |
| NH ₂ CF ₃ | 2.060 | 2.144 | 0.97 | 1656 | | | | |
| NH ₂ F | 2.386 | 2.386 | 0.86 | 1852 | | | | |
| | | | | | | | | |



Alternative localization ideas

Can break down when orbitals are delocalized

But it is also possible to study the electron pair without information from ELF, by projecting the EPLFN onto the orbital corresponding to the lone pair

$$\epsilon_{\textit{N}} = \frac{\langle \phi_{\rm lp} | {\rm EPLF} \textit{N} | \phi_{\rm lp} \rangle}{\langle \phi_{\rm lp} | \phi_{\rm lp} \rangle}$$

| EPLF projected on the lone pair orbital | | | | | | |
|---|-------------------------|-------------------------|------|--|--|--|
| Molecule | $\epsilon_0 	imes 10^2$ | $\epsilon_3 	imes 10^6$ | N | | | |
| TMA | 8.28 | 1.21 | 557 | | | |
| DMA | 8.34 | 1.21 | 514 | | | |
| MA | 8.48 | 1.96 | 659 | | | |
| NH_3 | 8.36 | 7.11 | 1299 | | | |
| NH ₂ CF ₃ | 8.32 | 7.19 | 1000 | | | |
| NH ₂ F | 8.95 | 5.41 | 1452 | | | |





EPLF3 in the ELF basin

A descriptor sensitive to the availability of the electron pair

Finally, the proposed descriptors from EPLFN, integrated over the basin defined by ELF, are presented.

| EPLF integrated on the ELF basin | | | | | | |
|----------------------------------|--------------|---------|-------------------------|---------|--|--|
| Molecule | ϵ_0 | d_0^c | $\epsilon_3 	imes 10^3$ | d_3^c | | |
| TMA | 5.16 | 2.35 | 2.89 | 3.52 | | |
| DMA | 5.59 | 2.38 | 3.10 | 3.45 | | |
| MA | 6.20 | 2.28 | 4.78 | 3.01 | | |
| NH_3 | 5.80 | 2.05 | 7.76 | 2.48 | | |
| NH ₂ CF ₃ | 4.28 | 2.07 | 6.15 | 2.70 | | |
| NH ₂ F | 5.10 | 1.93 | 7.02 | 2.43 | | |





Summary

- ϵ_3 and d_3^c are proposed as descriptors of the lone pair availability
- EPLF may not be the ultimate tool for analyzing electron pair localization
- EPLFN together with ELF provides a reliable, unambiguous description of the lone pair



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Outlook

- ullet Study the effect of more sophisticated trial wave functions $\psi_{
 m T}$
- Use forward-walking or reptation MC to obtain pure EPLFN estimator
- Currently applying to a set of 15 experimentally relevant molecules of 100 – 200 electrons.
- Explore EPLFN for the description of chemical bonds, radicals, and other kinds of lone pairs.



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Thanks!

José Alfredo Vázquez-Martínez (UNAM, México)

